Section 12 Practice

1) Write the forumula for the compounds made of the following ions:

a) calcium and iodine	b) sodium and oxygen	c) oxygen and magnesium
d) iron and oxygen	e) tin and bromine	f) potassium and suflur
g) copper and chloride	h) barium and iodine	i) alumium and oxygen
j) aluminum and chlorine		

2) Write the formula for the following compounds:

a) silver oxide	b) iron (III) sulfate	c) calcium nitrate
d) magnesium phosphate	e) copper (I) chloride	f) potassium carbonate
g) sodium bicarbonate	h) ammonium nitrate	i) barium hydroxide
j) ammonium oxide		

3) Name the following compounds:

$CuSO_4$	NaOH	Al_2O_3	$SrCO_3$
$SnCl_4$	V_2O_5	$Mg(ClO)_2$	$AlPO_4$
$LiNO_2$	$Pb(NO_3)_2$	$Zn(ClO_4)_2$	$CuCO_3$

4) Name or give the formula for the following hydrates:

$BaCl_2 \cdot 2H_2O$	calcium sulfate dihydrate
$LiCl \cdot H_2O$	sodium carbonate decahydrate
$MgSO_4 \cdot 7H_2O$	copper (I) cyanide pentahydrate
$Sr(NO_3)_2 \cdot 4H_2O$	zinc hydroxide dihydrate
$NH_4F \cdot 3H_2O$	aluminum nitrate tetrahydrate

5) Determine which compound has the higher lattice energy and explain why:

MgO vs. CaO	KCl vs. $CaCl_2$
Al_2O_3 vs. $AlCl_3$	NaBr vs. $MgBr_2$