1) Draw Lewis dot structures and name the following compounds that are exceptions to the octet rule. Include ALL lone pair electrons and formal charges.

ICl_5	**XeF_6**	**BrF_3**	**BH_3**	**PCl_5**
SCl_6	**H_3PO**	**IF_3**	**IF_5**	**SF_4**

2) Draw Lewis dot structures for the following ions, include formal charge

CH_3COO^-	CO_3^{-2}	CN^{-1}	N_3^{-1}	PO_4^{-3}
$CH_3NH_3^+$	***H_3O^+***	***OH^{-1}***	BH_4^{-1}	NO_2^{-1}

3) Draw ALL the resonance structures for the following molecules and circle most likely structure. Make sure to include formal charges.

CO_3^{2-}	CN^{-1}	N_3^{-1}	ClO_3^{-1}

4) Name and draw the following organic molecules:

a) CH_3CH_2OH	b) $CH_3CH_2OCH_2CH_3$	c) CH_3CH_2COOH	d) CH_3CH_2CHO
e) $CH_3CH_2CH_2CH_2CH_2OH$		f) $CH_3OCH_2CH_3$	g) $CH_3CH_2NH_2$
h) $CH_3CH_2CHCHCOOH$		i) $CH_3COCH_2CH_3$	j) $N(CH_3)_3$
k) $CH_3CH_2CCCH_2CHO$		l) $CH_3CH_2CH_2CH(OH)CH_3$	
m) $CH_3CH_2CH_2CH_2CH_2OH$		n) $CH_3CH(OH)CH_2CH_2CH_3$	
o) $CH_3CH_2CH(OH)CH_2CH_3$		p) $CH_3CH_2CH_2COOCH_2CH_3$	